## Acid reactions

## - overall and ionic equations

Hydrogen	Н+	Chloride	CI T
Sodium	Na <sup>+</sup>	Bromide	Br <sup>-</sup>
Silver	Ag+	Fluoride	F-
Potasssium	ΚŤ	Iodide	Ι-
Lithium	Li+	Hydroxide	OH-
Ammonium	NH₄ <sup>+</sup>	Nitrate	NO <sub>3</sub>
Barium	Ba <sup>2+</sup>	Oxide	02-
Calcium	Ca <sup>2+</sup>	Sulphide	S2-
Copper(II)	Cu <sup>2+</sup> Mg <sup>2+</sup>	Sulphate	SO <sub>4</sub> 2-
Magnesium	Mg <sup>2+</sup>	Carbonate	CO 2-
Zinc	Zn <b>z+</b>	Carbonate	CO <sub>3</sub> -
Lead	Pb <sup>2+</sup>	Hydrogencarbonate	
Iron(II)	Fe <sup>2+</sup>		HCO,
Iron(III)	Fe3+		11003
Aluminium	Al 3+		

Table 1
Valency of common ions

Reaction	
Sulphuric acid solution and	Overall :
aqueous sodium carbonate.	
	ionic :
Nitric acid solution and lithium	Overall :
metal	Overall:
metai	ionic :
Hydrochloric acid solution and	Overall :
magnesium sulphide powder	
	ionic :
Collaborate and anti-street	Overall :
Sulphuric acid solution and sodium oxide powder.	Overall :
Socialii Oxide powder.	ionic :
	Overall :
Nitric acid solution and	
aqueous calcium hydroxide.	ionic :
Hydrochloric acid solution and	Overall :
aqueous sodium sulphite.	Overall .
aqueous soulant sulprince.	ionic :
Sulphuric acid solution and	Overall :
solid sodium hydroxide	
	ionic :
Nitric acid solution and	Overall :
aqueous sodium carbonate.	Overan .
	ionic :

**Solutions**