

## Redox reactions - oxidation numbers

### Lesson 1

Redox reactions involve a transfer of electrons.

It is a two part process whereby electrons are lost and electrons are gained. Every redox reaction involves two processes, **oxidation** and **reduction**.

Oxidation is the process where a reactant loses electrons, this reactant is known as the **reductant**.

Reduction is the process where a reactant gains electrons, this reactant is known as the **oxidant**.

It is often difficult to know if a reactant has given or taken electrons. To overcome this we assign **oxidation numbers** to reactants and products. Oxidation occurs when the oxidation number of a reactant increases and reduction occurs when the oxidation number of a reactant. Oxidation numbers are assigned according to the following rules.

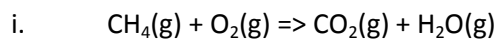
- i. The oxidation number of a species in its elemental form is 0.
- ii. The oxidation number on an elemental ion is the charge of the ion.  
Example the oxidation number of  $K^+$  is +1,  $Cl^-$  is -1 or  $O^{2-}$  is -2.
- iii. Hydrogen and oxygen found in compounds have fixed values.  
Example H is always +1 except in metal hydrides, such as  $CaH_2$  where it is -1.  
O is always -2 except in  $H_2O_2$  where it is -1 and in compounds with fluorine ( $OF_2$ ) where it is +2.
- iv. The sum of the oxidation numbers of each species in a compound is equal to the charge of that compound.  
Example  $SO_4^{2-}$  the sum of all the oxidation numbers is -2  
 $CO_2$  the sum of all the oxidation numbers is 0, since  $CO_2$  is a neutral compound.

Assign oxidation numbers to the underlined species.



Reactants that are considered reductants increase in oxidation number while those that are oxidants reduce in oxidation number. So a redox reaction has a reactant that increases in oxidation number and one that decreases in oxidation number.

For each of the following reactions identify the redox reactions and give the reductant and oxidant in each



Find the oxidation of each element in the reactants. C = -4, H = +1, O = 0

Find the oxidation of each element in the products. C = +4, H = +1, O = -2

Reductant is the carbon in  $\text{CH}_4$

Oxidant is the oxygen in  $\text{O}_2$

