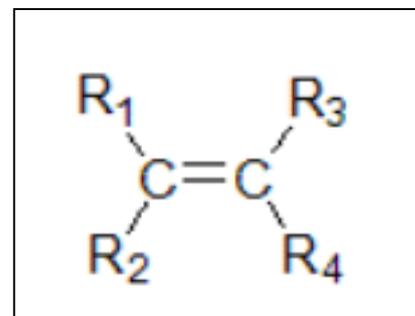


## Lesson 4 Naming cis and trans isomers and writing their semistructural formulae

[Click](#) to revise cis and trans isomers

[Click](#) to revise structural and semistructural formulae

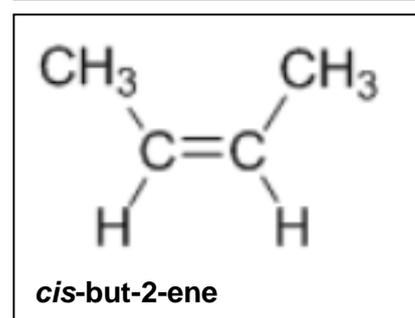
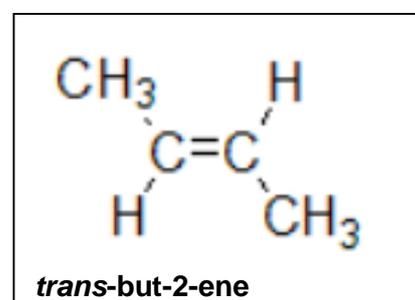
An alkene can exist as *cis* or *trans* isomers *only* if  $R_1$  is not equal to  $R_2$  and  $R_3$  is not equal to  $R_4$



*cis*- if the two alkyl groups, R-, are on the same side of the C=C

*trans*- if the two alkyl groups, R-, are on opposite sides of the C=C.

the terms ***cis*** and ***trans*** are inserted into the name as prefixes.

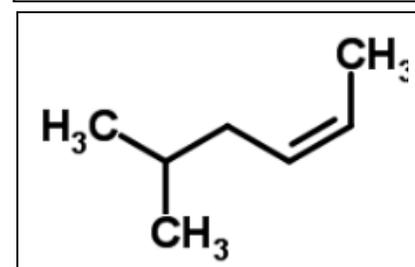


### Example 1

Name the molecule shown on the right.

This molecule is a *cis* isomer of 5-methylhex-2-ene

We write the name with the *cis* prefix ***cis*-5-methylhex-2-ene**. The semi-structural formula is given as ***CH<sub>3</sub>CH(CH<sub>3</sub>)CH<sub>2</sub>CHCHCH<sub>3</sub>***.

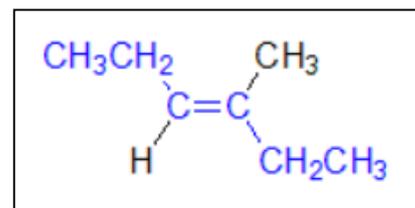


### Example 2

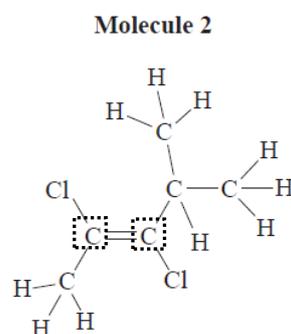
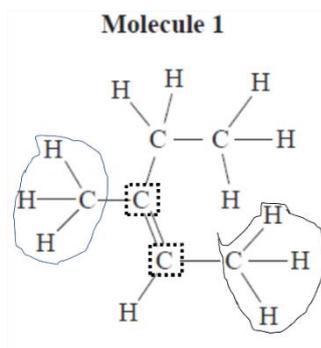
Name the molecule shown on the right.

This molecule is a *trans* isomer of 3-methylhex-3-ene

We write the name with the *trans* prefix ***trans*-3-methylhex-3-ene**. The semi-structural formula is given as ***CH<sub>3</sub>CH<sub>2</sub>CHC(CH<sub>3</sub>)CH<sub>2</sub>CH<sub>3</sub>***

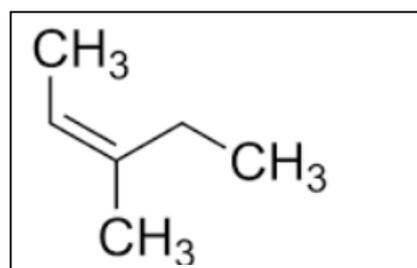


Below is a recommendation from the Chief assessor in the 2021 assessment response. Follow the Assessors comments when answering the questions below. "Geometric – *cis/trans* – isomers of alkenes can exist when the same atoms/groups are attached at both ends of the C=C double bond. At VCE level students should be aware that when these atoms/groups are on the same side of the C=C, the compound is the *cis* isomer, and when they are on different sides, the compound is the *trans* isomer. In Molecule 1 there are methyl, CH<sub>3</sub>-, groups on different sides of C=C, so on that basis it would be classified a *trans* structure (i.e. *trans*-3-methylpent-2-ene). Molecule 2 there are chloro, Cl-, groups on different sides of the molecule, so on that basis it would also be classified as a *trans* structure (i.e. *trans*-2,3-dichloro-4-methylpent-2-ene)."

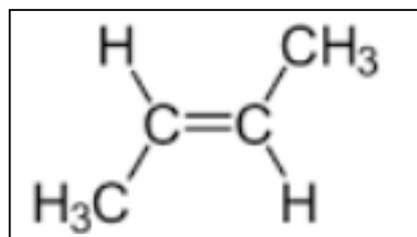


1) Name the molecules shown on the right and give their semi-structural formulae

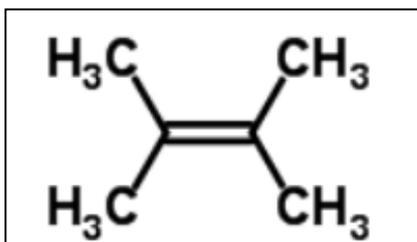
a)



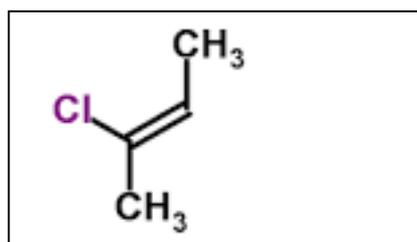
b)



c)



d)



e)

