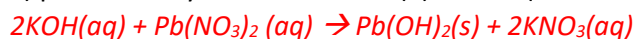


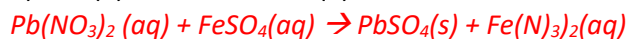
# Solutions

Soluble Ionic Compounds		Important Exceptions
Compounds containing	NO <sub>3</sub> <sup>-</sup>	None
	C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>	None
	Cl <sup>-</sup>	Compounds of Ag <sup>+</sup> , Hg <sup>2+</sup> , and Pb <sup>2+</sup>
	Br <sup>-</sup>	Compounds of Ag <sup>+</sup> , Hg <sup>2+</sup> , and Pb <sup>2+</sup>
	I <sup>-</sup>	Compounds of Ag <sup>+</sup> , Hg <sup>2+</sup> , and Pb <sup>2+</sup>
	SO <sub>4</sub> <sup>2-</sup>	Compounds of Sr <sup>2+</sup> , Ba <sup>2+</sup> , Hg <sup>2+</sup> , and Pb <sup>2+</sup>
Insoluble Ionic Compounds		Important Exceptions
Compounds containing	S <sup>2-</sup>	Compounds of NH <sub>4</sub> <sup>+</sup> , the alkali metal cations, and Ca <sup>2+</sup> , Sr <sup>2+</sup> , and Ba <sup>2+</sup>
	CO <sub>3</sub> <sup>2-</sup>	Compounds of NH <sub>4</sub> <sup>+</sup> and the alkali metal cations
	PO <sub>4</sub> <sup>3-</sup>	Compounds of NH <sub>4</sub> <sup>+</sup> and the alkali metal cations
	OH <sup>-</sup>	Compounds of the alkali metal cations, and Ca <sup>2+</sup> , Sr <sup>2+</sup> , and Ba <sup>2+</sup>

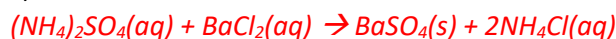
a) potassium hydroxide and lead(II) nitrate (Click for the solution to a)



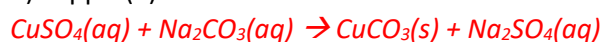
b) lead(II) nitrate and iron(II) sulfate



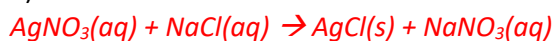
c) ammonium sulfate and barium chloride



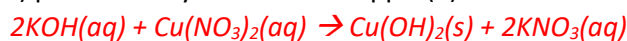
d) copper(II) sulfate and sodium carbonate



e) silver nitrate and sodium chloride



f) potassium hydroxide and copper(II) nitrate

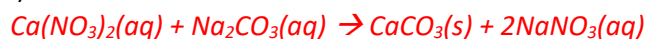


g) sodium hydroxide and iron(III) nitrate (Click for the solution to g)

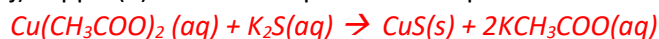
h) aluminium chloride and sodium hydroxide



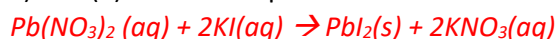
i) calcium nitrate and sodium carbonate



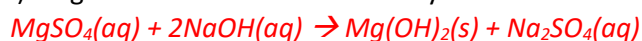
j) copper(II) acetate and potassium sulphide



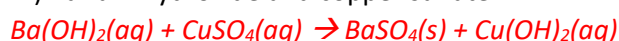
k) lead(II) nitrate and potassium iodide



l) magnesium sulfate and sodium hydroxide



m) Barium hydroxide and copper sulfate



n) Calcium sulfide and aluminium sulfate



*Two precipitates form.*